**THE PLAN**

**of General Chemistry course laboratory practical classes**

**for the students of general medicine faculty**

**and foreign students faculty**

**in the I semester of 2018-2019 academic year.**

**(practical class duration — 2 hours)**

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| № | Topic |
| 1 | Introduction in practical. Introductory lesson. |
| 2 | The theory of aqueous solutions. General characteristics of solutions. Ways of expression for aqueous solutions composition. |
| 3 | Chemical equivalent. Equivalents law. |
| 4 | Basics of titrimetric analysis. |
| 5 | Acid-base titration. |
| 6 | The theory of redox reactions. Methods of redox reaction balancing. Bases of redoximetric titration. |
| 7 | Colligative properties of diluted solutions of nonelectrolytes and electrolytes. |
| 8 | Equilibrium biochemical processes. Buffer solutions. |
| 9 | Equilibrium in the solutions of Complex (coordinate) compounds. |
| 10 | Basics of chemical kinetics. Catalysis. Chemical equilibrium. |
| 11 | Equilibrial electrode processes. Potentiometry. |
| 12 | Equilibrial electrode processes. Measurements of redox potential. |
| 13 | Physico-chemistry of surface phenomena. Adsorption on the mobile phase border of partition. |
| 14 | Physico-chemistry of surface phenomena. Adsorption on the stationary phase border of partition. |
| 15 | Physico-chemistry of disperse systems. Properties of colloidal systems. |
| 16 | Sols stability and coagulation. |
| 17 | Physico-chemistry of biopolymers solutions. Protective action of HMC. |
| 18 | Pre-session practical. |

Head of General and Bioorganic

Chemistry Department,

associate professor (docent) V.V. Boltromeyuk