

**THE PLAN**  
**of General Chemistry course laboratory practical classes**  
**for the students of general medicine faculty**  
**and foreign students faculty**  
**in the I semester of 2020-2021 - academic year.**  
**(practical class duration — 2 hours)**

№	Topic
1	Introduction in practical. Introductory lesson.
2	The theory of aqueous solutions. General characteristics of solutions. Ways of expression for aqueous solutions composition.
3	Chemical equivalent. Equivalent law.
4	Basics of titrimetric analysis.
5	Acid-base titration.
6	The theory of redox reactions. Methods of redox reaction balancing. Bases of redoximetric titration.
7	Colligative properties of diluted solutions of nonelectrolytes and electrolytes.
8	Equilibrium biochemical processes. Buffer solutions.
9	Equilibrium in the solutions of Complex (coordinate) compounds.
10	Basics of chemical kinetics. Catalysis. Chemical equilibrium.
11	Equilibrium electrode processes. Potentiometry.
12	Equilibrium electrode processes. Measurements of redox potential.
13	Physico-chemistry of surface phenomena. Adsorption on the mobile phase border of partition.
14	Physico-chemistry of surface phenomena. Adsorption on the stationary phase border of partition.
15	Physico-chemistry of disperse systems. Properties of colloidal systems.
16	Sols stability and coagulation of sols.
17	Physico-chemistry of biopolymers solutions. Protective action of HMC.
18	Chemical bonds, valency, oxidation number.

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associate professor (docent)



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